

Acids and bases review sheet Answer key

1. List the differences between an acid and a base. Your description should include such things as whether both are electrolytes, which contains a hydrogen ion and which includes a hydroxide ion, and you should also include two examples of an acid and two examples of a base that are present in your household.

| Acid | base |
|------------------------------|-------------------------|
| Conducts electricity | conducts electricity |
| Contains hydrogen ion | contains hydroxide ion |
| Proton donor | proton acceptor |
| Turns litmus paper red | turns litmus paper blue |
| React with most metals | releases hydroxide ion |
| Releases hydronium | Ex: Windex, |
| Ex: lemon | bleach |
| Vinegar | soap |
| Battery acid (sulfuric acid) | |

2. Listed below are some of the properties of acids and bases. Fill in the blanks with the appropriate word, **acids** or **bases**:

- A** produce hydrogen ions (H^+) in solution
- B** have a bitter taste
- B** produce hydroxide ions (OH^-) ions
- A** react with bases to form salts
- A** have a sour taste
- B** have a slippery, soapy feel
- B** react with acids to form salts
- A** react with many metals

3. Write an equation for the dissociation of the following **acids** in water: follow example

- $HClO_4 \rightarrow H^+ + ClO_4^-$
- $H_2SO_4 \rightarrow 2H^+ + SO_4^{2-}$
- $HC_2H_3O_2 \rightarrow H^+ + C_2H_3O_2^-$
- $H_2S \rightarrow 2H^+ + S^{2-}$
- $HCl \rightarrow H^+ + Cl^-$
- $HNO_3 \rightarrow H^+ + NO_3^-$

4. Write an equation for the dissociation (ionization) of the following **bases** in water: (follow example)

- $Mg(OH)_2 \rightarrow Mg^{2+} + 2OH^-$
- $NaOH \rightarrow Na^+ + OH^-$
- $Ba(OH)_2 \rightarrow Ba^{2+} + 2(OH^-)$
- $KOH \rightarrow K^+ + OH^-$
- $Mg(OH)_2 \rightarrow Mg^{2+} + 2(OH^-)$
- $LiOH \rightarrow Li^+ + OH^-$

5) A. What type of reaction is a **neutralization** reaction? **Double replacement**

B. Write the net equation for neutralization:



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