$q = m \cdot c \cdot \Delta T$

Thermal Energy & Heat Unit Study Guide

- Part 1. For each of the following scenarios:
 - a) Identify as vaporization, condensation, melting, or freezing
 - b) Write the change in states [example: gas → liquid]
 - c) Write whether the process absorbs or releases thermal energy



1. High altitude rain turns to hail.

a) _____ c) ____

- Towering cumulonimbus clouds form in the sky.

a) ______ c)___

- A shallow puddle dries up over a couple day span.

Part II. Specific Heat Calculations Refer to the equation and symbol legend above. Hint use the triangle method. Show your work.

1. A 110 g sample (m) of tin was heated from 23.2 °C to 39.8 °C. What was the change in thermal energy? Solve for the heat energy (g). Tin has a specific heat (c) value of 0.2274 J/g-°C*.

c =

2. A scientist heated bundled strands of copper wire from 21.0 °C to 82.0 °C. Copper has a specific heat (c) capacity of 0.3845 J/g-°C*. If the change in heat energy (q) was 192 J. what was the mass (m) of the bundled strands of copper wire?

Thermal Energy And Heat Study Guide Answers

Clifford Lane

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