

Chapter 17 Thermochemistry Study Guide

17.1 – 17.2 Thermochemical Equations

1. Make the following conversions:
 - a. 444 cal to joules = $1.86 \times 10^3 \text{ J}$
 - b. 1.8 kJ to joules = $1.8 \times 10^3 \text{ J}$
 - c. 0.45 kJ to calories = $1.1 \times 10^2 \text{ cal}$
2. Classify each of these processes as endothermic or exothermic:
 - a. condensing steam – **exo**
 - b. burning alcohol – **exo**
 - c. evaporating alcohol – **endo**
 - d. baking a potato – **endo**
3. The specific heat capacity for silver is $0.24 \text{ J/g}^\circ\text{C}$. Calculate the energy required to raise the temperature of 150.0 g Ag from 273 K to 298 K. Calculate the molar heat capacity of silver.

$$\text{Energy: } q = (150.0)(.24)(298-273) = 9.0 \times 10^2 \text{ J}$$

$$\text{Molar heat capacity: } (0.24 \text{ J/g}^\circ\text{C})(107.87 \text{ g/mol Ag}) = 26 \text{ J/mol}$$

4. It takes 585 J of energy to raise the temperature of 125.6 g Hg from 20.0°C to 53.5°C . Calculate the specific heat capacity and the molar heat capacity of Hg.

$$\text{Specific heat capacity: } C = q/m\Delta t = 585/(125.6)(53.5-20) = 0.139 \text{ J/g}^\circ\text{C}$$

$$\text{Molar heat capacity: } (0.139 \text{ J/g}^\circ\text{C})(200.59 \text{ g/mol Hg}) = 27.9 \text{ J/mol}$$

5. A 46.2-g sample of copper is heated to 95.4°C and then placed in a calorimeter containing 75.0 g water at 19.6°C . The equilibrium temperature in the calorimeter is 21.8°C . Calculate the specific heat capacity of copper; assuming that all the heat lost by the copper is gained by the water.

$$C = 0.203 \text{ J/g}^\circ\text{C}$$

6. A 15.0-g sample of nickel metal is heated to 100.0°C and dropped into 55.0 g of water, initially at 23.0°C . Assuming that all the heat lost by the nickel is absorbed by the water; calculate the final temperature of the nickel and the water. The specific heat of nickel is $0.444 \text{ J/g}^\circ\text{C}$.

$$\begin{array}{rcl} 15.0(.444)(100.0 - T_f) & = & 55.0(4.184)(T_f - 23.0) \\ 6.66T_f & = & 230.12T_f - 5292.76 \\ 5958.76 & = & 236.78 T_f \\ \mathbf{25.2^\circ\text{C}} & = & T_f \end{array}$$

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of the year award i would like to nominate mary jane for the employee of the year award mary is a dedicated punctual

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web mention the award for which the candidate is nominated in the endorsement mention the kind of relationship you share with the candidate and since how long you know him her state the reason for which the award is given mention the qualities and achievements of the candidate due to which he she is nominated

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web jul 19 2023 1 list positive employee contributions before you draft your letter consider listing the employee s positive contributions to the company think of specific actions behaviors or accomplishments you d like to acknowledge or reinforce

being specific can help the employee feel valued because it shows you took the time to notice their hard work

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