Unit 6: Thermodynamics Study Guide and Practice Test Topics for this exam:

- 6.1 Endothermic and Exothermic Processes
- 6.2 Energy Diagrams
- 6.3 Heat Transfer and Thermal equilibrium
- 6.4 Heat Capacity and Calorimetry

- 6.5 Energy of Phase Changes
- 6.6 Enthalpy of Reaction
- 6.7 Bond Enthalpy
- 6.8 Enthalpy of Formation
- 6.9 Hess's Law

Vocabulary

- System
- Surroundings
- Universe
- State Function
- Energy
- Enthalpy
- Entropy
- Gibb's Free Energy

- Endothermic
- Exothermic
- Kinetic Energy
- Potential Energy
- Specific Heat
- Heat Capacity
- Work
- Joule

- Calorie
- Standard State
- Hess's law
- Enthalpy of
 - Formation
- Enthalpy of Reaction

Practice Test- Multiple Choice

1) Which of the following reactions have a negative entropy change?

- I. 2 $H_2(g) + O_2(g) \rightarrow 2 H_2O(1)$
- II. 2 NH₃(g) \rightarrow N₂(g) + 3 H₂(g)
- III. $Ca(s) + Cl_2(g) \rightarrow CaCl_2(s)$
- (A) II only
- (B) I only

- (C) I, II, and III
- (D) I and III

2) A certain reaction is non-spontaneous under standard conditions, but becomes spontaneous at higher temperatures. What conclusions may be drawn under standard conditions?

- (A) $\Delta H < 0$, $\Delta S > 0$, and $\Delta G > 0$
- (B) $\Delta H > 0$, $\Delta S < 0$, and $\Delta G > 0$

- (C) $\Delta H > 0$, $\Delta S > 0$, and $\Delta G > 0$
- (D) $\Delta H < 0$, $\Delta S < 0$, and $\Delta G > 0$

3) The value of ΔHe for the reaction below is -126 kJ. How much heat (in kJ) is released when 2.00 mol of NaOH is formed in the reaction?
2Na₂O₂ (s) + 2H₂O (l) → 4NaOH (s) + O₂ (g)

A) 252

B) -126

C) 7.8

D) 63

E) 3.9

3 $C_2H_2(g) \rightarrow C_6H_6(g)$ What is the standard enthalpy change, ΔH° , for the reaction represented above? (ΔHf° of $C_2H_2(g)$ is 230 kJ mol-l; ΔHf° of $C_6H_6(g)$ is 83 kJ mol-1.) (A) -607 kJ (B) -147 kJ (C) -19 kJ (D) +19 kJ

5)
$$CH_4(g) + 2 O_2(g) \rightarrow CO(g) + 2 H_2O(l); \Delta Hr^{\circ} = -889.1 \text{ kJ}$$

 $\Delta Hf^{\circ} H_2O(l) = -285.8 \text{ kJ} / \text{mole}$

Thermodynamics Study Guide Answers

SJ Ball

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