

## Thermochemistry Review

1. A gas absorbs 10.0 J of heat and then performs 15.2 J of work. The change in internal energy of the gas is:  
 a) 10.0 J  
 b) -15.2 J  
 c) 5.2 J  
 d) 25.2 J  
 (e) -5.2 J
2. Which of the following are state functions?  
 a) work  
 b) enthalpy  
 c) internal energy  
 (d) two of the above  
 e) all of the above
3. A 40.0 g sample of a metal is heated to 99.3 °C and then placed in a calorimeter containing 120.0 g of water at 21.8 °C. The final temperature of the water is 24.5 °C. Which metal was used? ? Given that the specific heat of water = 4.184 J/g °C.  
 a) Al, specific heat = 0.89 J/g °C  
 (b) Fe, specific heat = 0.45 J/g °C  
 c) Ag, specific heat = 0.24 J/g °C  
 d) Cu, specific heat = 0.20 J/g °C  
 e) Pb, specific heat = 0.14 J/g °C

4. What is the specific heat of mercury if it requires 167 J to change the temperature of 15.0 g of mercury from 25.0 °C to 33.0 °C

$$167 \text{ J} = m \cdot c \cdot \Delta T \quad 167 = 15 \cdot c \cdot (33 - 25)$$

$$167 = 15 \cdot c \cdot 8$$

$$167 = 120 \cdot c$$

$$c = \frac{167}{120} = 1.39 \text{ J/g}^\circ\text{C}$$

?

cal from  
500 J  
137

5. Based on the reaction below, what energy change occurs when 1.20 moles of HCl is produced?



$$1.2 \text{ moles} \times \frac{-434 \text{ kJ}}{4 \text{ moles HCl}} = -130.2 \text{ kJ}$$

$$= -130.2 \text{ kJ}$$

6. How much heat is required to raise the temperature of a 6.21 g sample of iron from 25.0 °C to 79.8 °C? Given that the specific heat of iron = 0.450 J/g °C.

$$X = 0.450 \cdot 6.21 \cdot 54.8$$

$$+ 153.4 \text{ J}$$

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